



Revision Techniques

Here are some hints and tips

Different ways to revise:

- **Reading**
- **Making notes**
- **Checklists**
- **RAG rating**
- **Revision cards**
- **Mnemonics**
- **Association**
- **Record Information**
- **Example questions and answers**
- **Past paper questions**
- **'Revise, test, analyse'**
- **Mind Map**

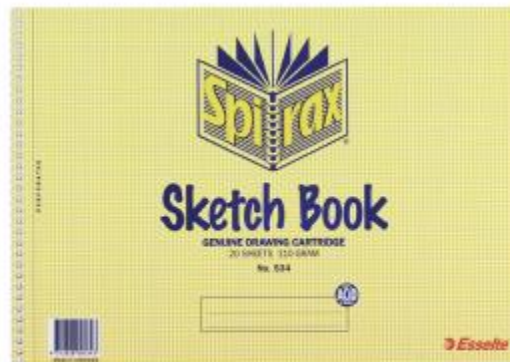
What type of learner are you?

Visual learner	Auditory learner	Reader and writer	Active learner
<ul style="list-style-type: none">• rewrite your notes as mind-maps• use colour to highlight important facts• draw diagrams and sketches to help you remember points	<ul style="list-style-type: none">• read your notes aloud• record yourself reading key points aloud to listen back to• revise in groups• try singing the key points and associate them with a tune	<ul style="list-style-type: none">• copy out your notes again• read over old notes• rewrite key points using different phrases• try to write key points from memory, and test yourself	<ul style="list-style-type: none">• move around the room or carry out an activity as you revise• mentally test yourself while you are exercising to see what you can remember

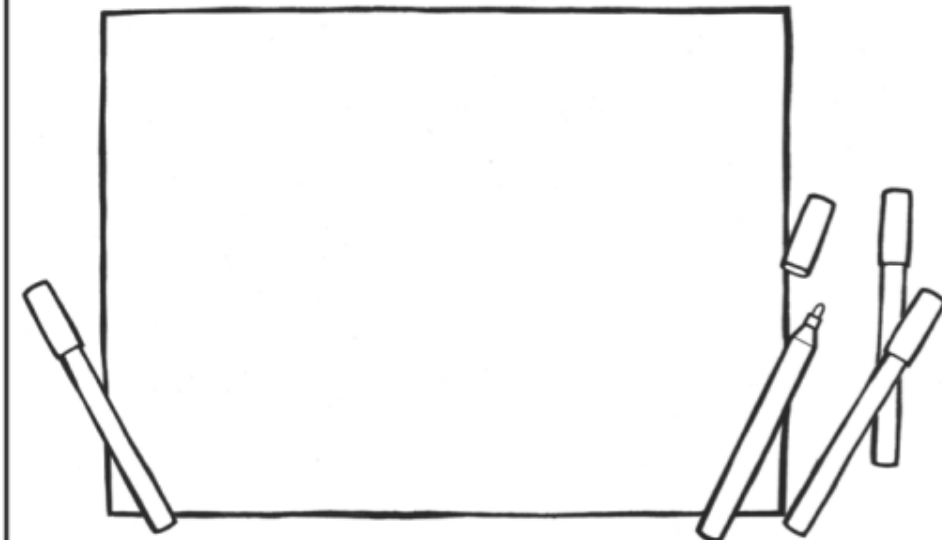
Mind Map

- Start at the centre with the page with a circle or square ('central node') indicating the main topic.
- The main ideas that make up the topic should be represented by thick branches that emanate from the central node. Your branches should have one word or two for the purpose of clarity. Think of each branch as a heading in an essay or a book.
- Create smaller sub-branches which extend out from every branch. Think of these as sub-headings.
- At the end of each branch, write out one key word or concept. This will make it easier to remember key concepts during revision.
- Find images or pictures which illustrate your ideas and paste them onto key areas in your diagram.
- Try to find connections between the various branches and sub-branches; this is a crucial aspect of critical thinking.
- Use colours and pictures

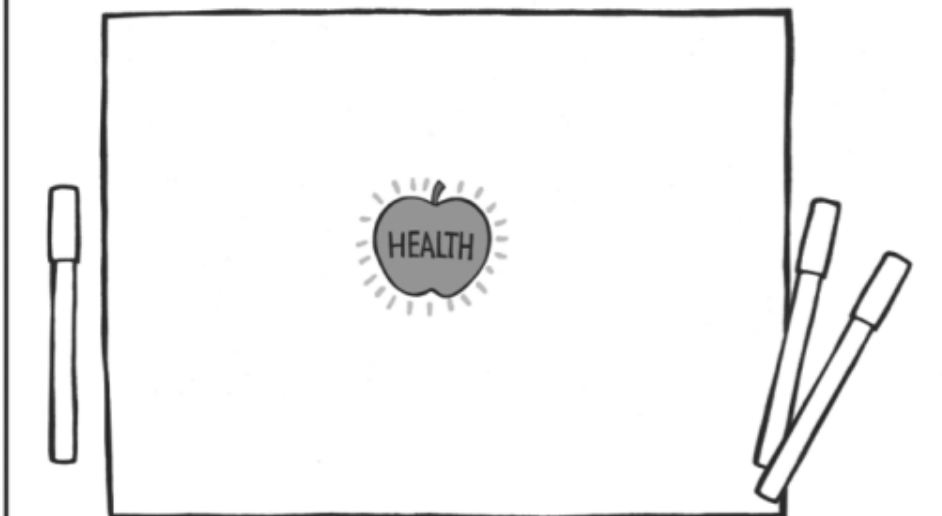
Set yourself up ...



STEP 1 : YOU WILL NEED THE FOLLOWING MATERIALS : COLOURED PENS AND BLANK PAPER (NO LINES).



STEP 2 : MAKE SURE YOUR PAGE IS TURNED SIDEWAYS (LANDSCAPE STYLE), THINK OF YOUR TOPIC FOR A MOMENT AND THEN DRAW THE FIRST IMAGE THAT JUMPS TO MIND IN THE CENTRE OF THE PAGE.



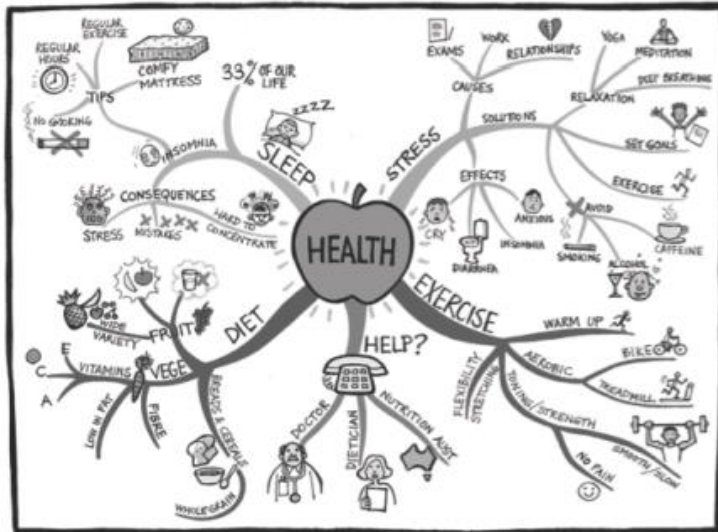
STEP 3 : BRANCH OFF. THINK OF THE BRANCH AS BEING LIKE A CHAPTER IN A BOOK. IT'S ONE OF YOUR MAIN IDEAS.



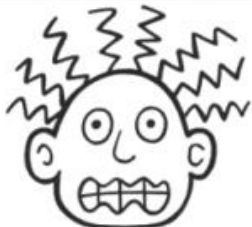
STEP 4 : NEXT CREATE YOUR SUB-BRANCHES. THESE BRANCHES ARE LIKE THE SUB-HEADINGS WITHIN THE CHAPTER OF A BOOK. FROM THE ENDS OF THESE BRANCHES, YOU CAN CREATE MORE BRANCHES. (WHAT YOU ARE DOING IS CREATING ASSOCIATIONS BETWEEN DIFFERENT IDEAS).



STEP 5 : DRAW PICTURES. IF I TOLD YOU "DON'T THINK ABOUT MONKEYS", WHAT WOULD COME TO MIND? THE PICTURE OF MONKEYS! MY POINT IS, WE THINK IN PICTURES, NOT IN WORDS. THE MORE EXAGGERATED YOU MAKE THE IMAGES ON YOUR MIND MAP, THE EASIER IT WILL BE FOR YOU TO RECALL THEM.



EXAMPLES OF EXAGGERATED IMAGES



STRESS



INSOMNIA



HARD TO
CONCENTRATE



CRY



SET GOALS



ALCOHOL

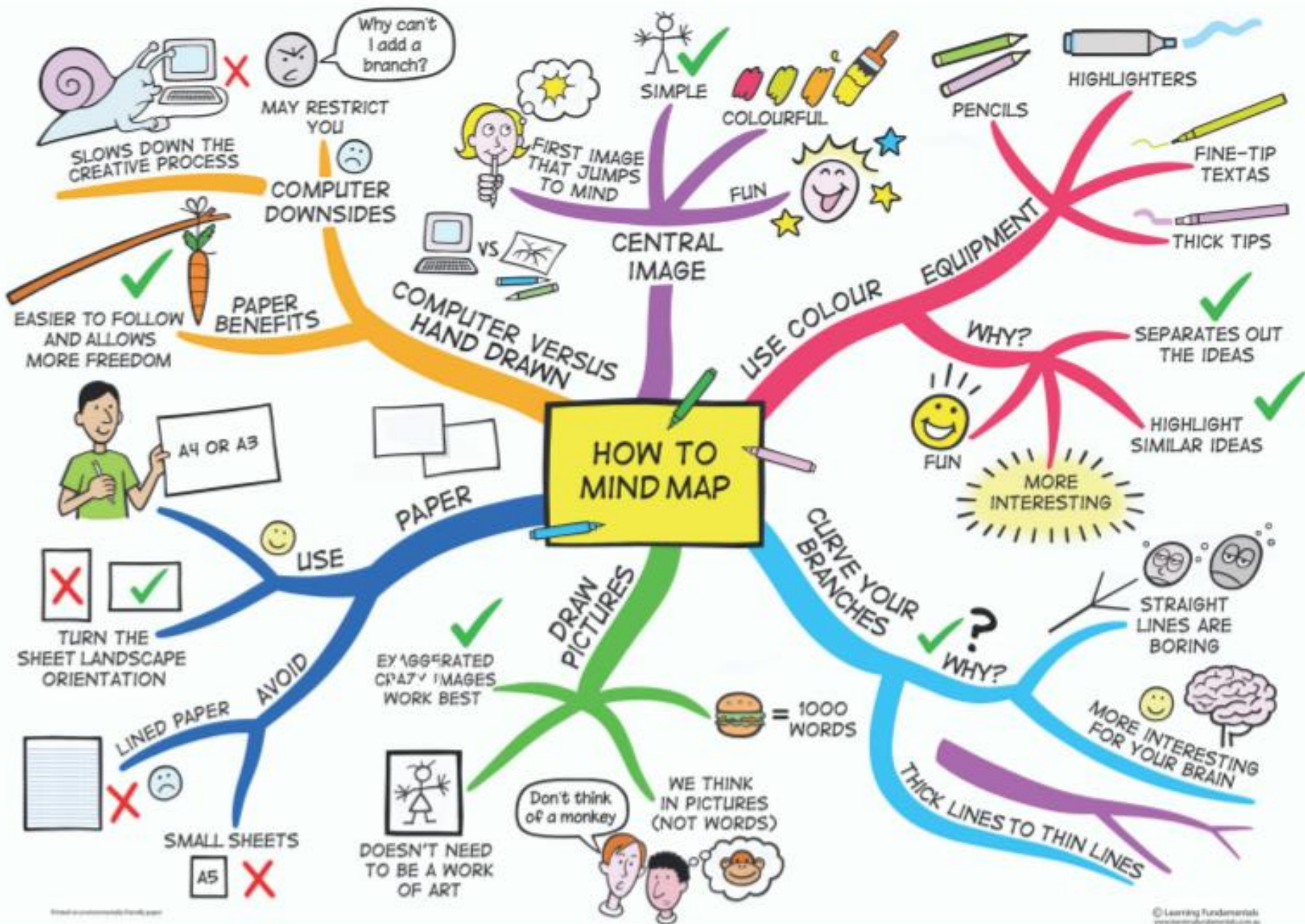
Only mind map the 'good' stuff ...

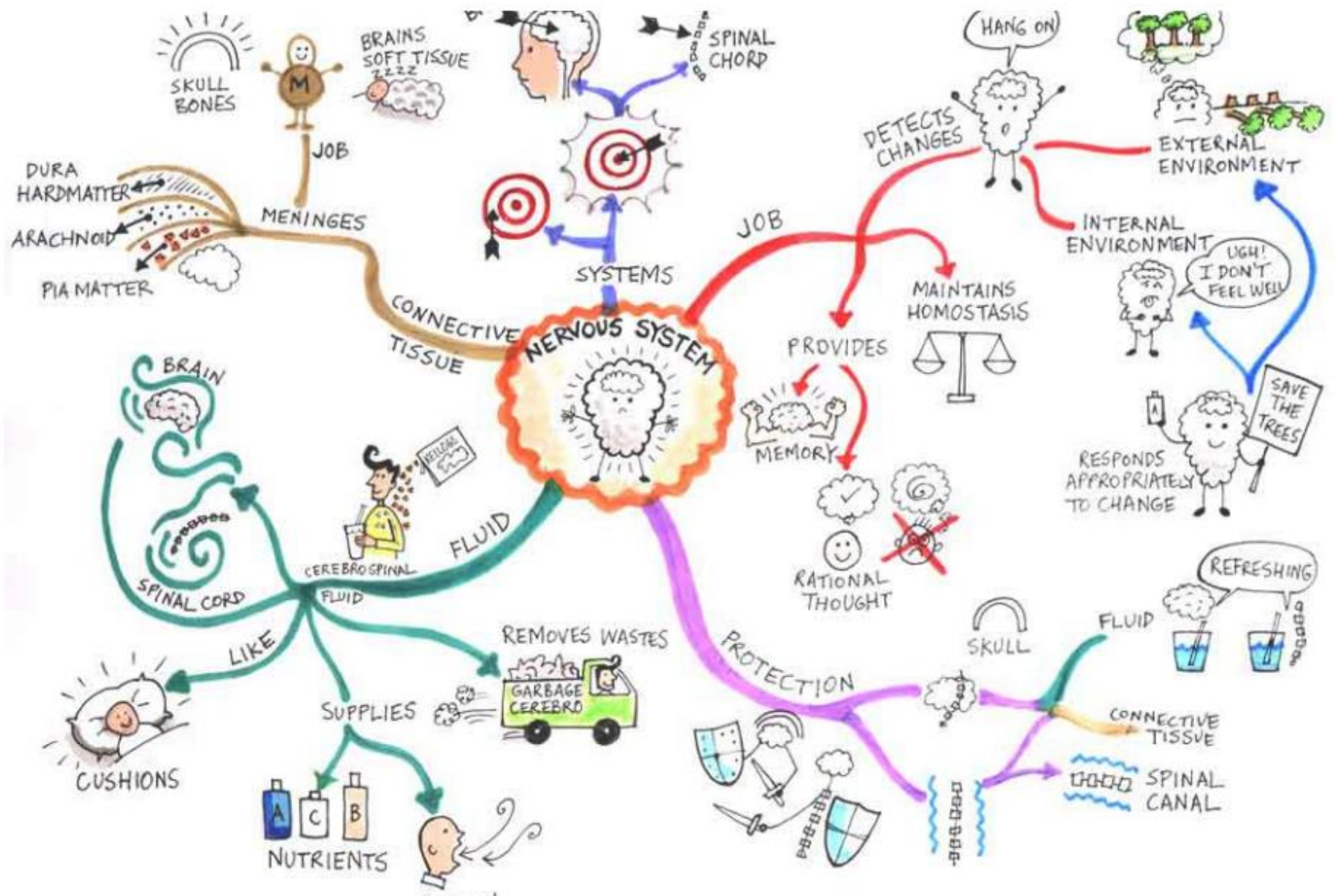
In my experience mind maps work best when you're trying to understand complex ideas. Some information you need to absorb will be straightforward and can be put straight onto a flashcard (you may not need to mind map it). Other information won't be so relevant.

The bottom line is not everything needs to be mind mapped. As you read your book and mind map the information, ask yourself:

"Do I really need to know this?"

If the answer is no, then don't bother mind mapping it.





Where is the Earth's Water?

Seas
Lakes
Sky

Water Cycle

INPUTS - Precipitation

OUTPUTS

- Evaporation
- Transpiration
- River channel discharge

THROUGHFLOW

- Interception
- Infiltration
- Percolation
- Run-off
- Groundwater

WATER & THE WATER CYCLE

ADVANTAGES

- 1...
- 2...
- 3...
- 4...

AQUIFERS

Must be permeable & porous

Problems of over abstraction

SOURCES OF WATER

RESERVOIRS

How do they affect the river downstream?

How do they affect the microclimate?

UK

Demand for Water

Future Demand - What? Reasons?

USES

Industrial
Domestic
Agricultural

DESALINATION

Reverse Osmosis
Distillation

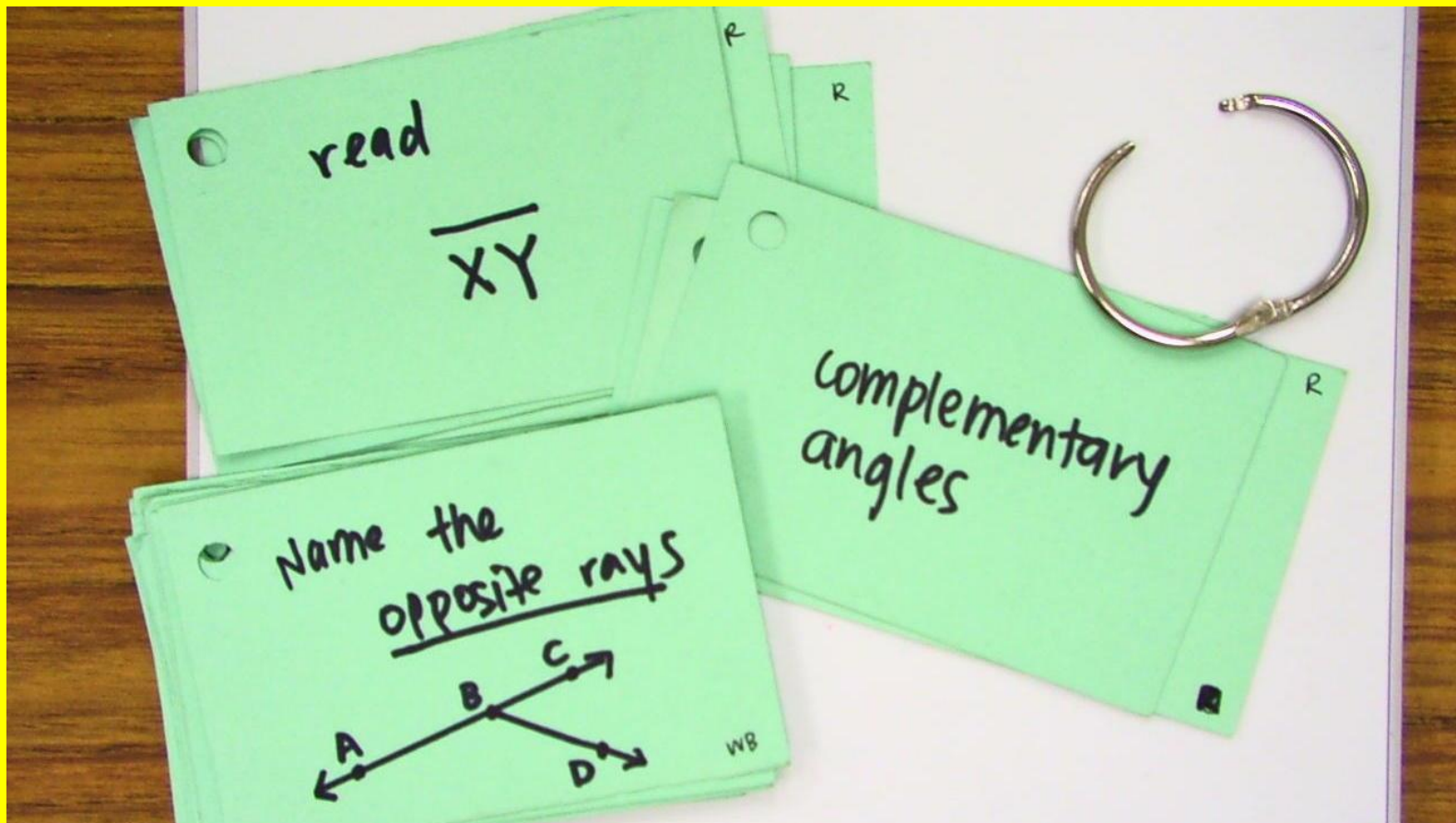
SEWAGE TREATMENT

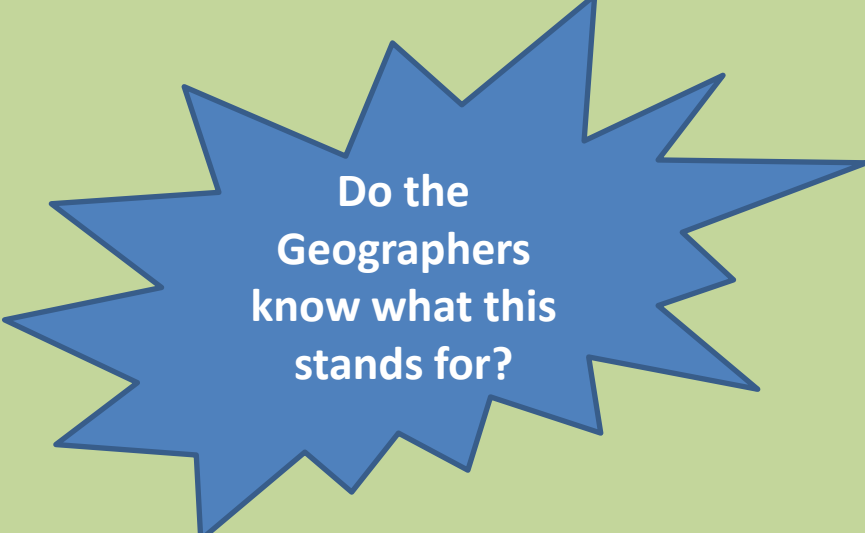
- Grit Trap
- Screens
- Settling Tanks
- Phosphate Stripping
- Sludge tanks aerated
- Anabolic Sludge Digestion

Used to make m...

Flash cards

- Make them and then test yourself





Do the
Geographers
know what this
stands for?

- C
- C
- A
- S
- H

Mnemonics

Can be a helpful way to memorise facts. The first letter is used to create a phrase or word that you can more easily remember.

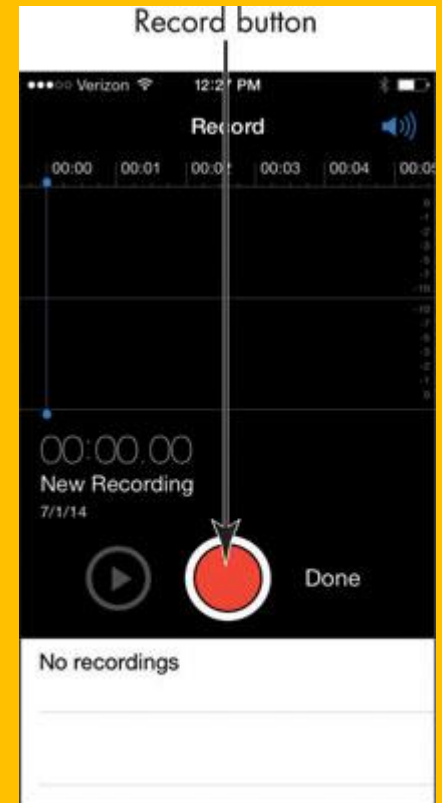
For example, Richard Of York Gave Battle In Vain (colours of the rainbow in order: Red, Orange, Yellow, Green, Blue, Indigo, Violet).

Association

Some students find it helpful to associate sounds or images with the information they are learning or to use notes / mind maps as a visual memory trigger.

Record Information ...

.... That you need to remember and then you can play it back to yourself



A3 sheet

Separation Techniques Revision - Activate

Explain in detail the differences between mixtures and compounds. – p72

With the aid of a diagram, show how a substance dissolves into a liquid – p74

What does solubility mean? – p76

What things can affect solubility? – p77

Explain the process and the result of a chromatography test on a felt tip pen. – p82

What is a saturated solution? – p76

Define these words:

Solvent:

Solute:

Dissolve:

Soluble:

Filtrate:

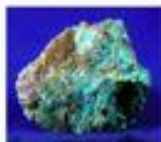
Residue:

Using a diagram, show how filtration works. – p78

Describe what distillation is and how it works. – p81

Explain how people get salt from salt water. – p80

Ores are naturally occurring rocks that contain metal or metal compounds in sufficient amounts to make it worthwhile (economically) extracting them



Reactivity of metals

potassium	electrolysis of a molten compound
sodium	
calcium	
magnesium	
aluminium	
zinc	heat an ore with carbon
iron	
tin	
lead	
copper	found as the uncombined element
silver	
gold	
platinum	

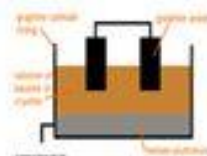
most reactive (hard to extract)

least reactive (easy to extract)

Extracting more reactive metals

(electrolysis)

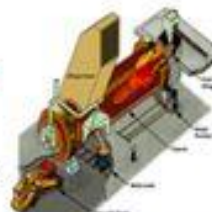
- Electrolysis involves passing an electric current through a metal compound
- This breaks it down into metal and non metal elements
- Aluminium is extracted in this way. *Why?*
- This uses large amounts of energy, making these metals expensive



Molten Aluminium oxide $\xrightarrow{\text{Electricity}}$ molten aluminium + oxygen gas

How extraction works

- The ore is crushed to remove rock with little metal
- The metal oxide has the oxygen removed
- This is called a reduction reaction
- Reduction is achieved by heating the metal oxide with carbon



Iron oxide + Carbon $\xrightarrow{\text{Heat}}$ Iron + carbon dioxide

Extraction of copper from copper oxide by reduction

The copper oxide is **REDUCED**, which means it **LOSES oxygen**.

Copper oxide + carbon \rightarrow Copper + carbon dioxide



An Oxidation Reaction



When iron corrodes (rusts) we say an oxidation reaction is happening

Oxidation is the addition of oxygen to something

Gold!

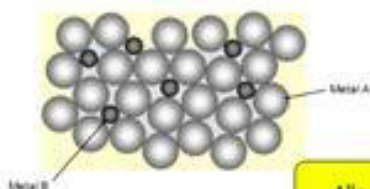
Jewellery made from pure gold is **very soft**. Gold jewellery is usually made from alloys, in which gold is mixed with other metals



Percentage gold	Percentage other metals	Carats	Precious metal per 1000g
100%	0%	24	1000
91.7%	8.3%	22	917
75%	25%	18	750
50%	50%	12	500
27.8%	72.2%	9	278

By mixing two or more metals together the regular arrangement of the atoms is disrupted

This prevents the atoms forming layers and makes it harder for the atoms to slide over each other. The alloy is **stronger** than the pure metals



Alloy

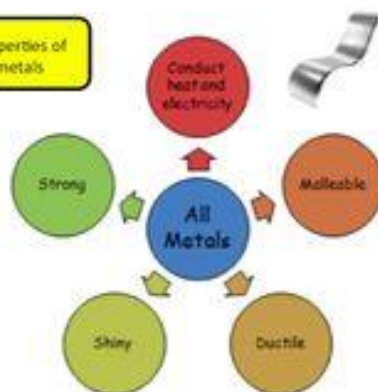
Stainless steel is an alloy that contains other metals such as **Nickel** and **Chromium**.

Stainless steel has the strength of steel combined with the corrosion resistance of nickel and the shininess of chromium.



Shape memory alloys (e.g. Stent to keep a blocked blood vessel open)

Properties of metals



Reduction and Oxidation

Some examples of reduction:

Aluminium + iron oxide $\xrightarrow{\text{heat}}$ aluminium oxide + iron

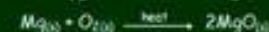


Lead oxide + carbon $\xrightarrow{\text{heat}}$ lead + carbon dioxide



An example of oxidation:

Magnesium + oxygen $\xrightarrow{\text{heat}}$ magnesium oxide



Note-taking planner

[illegible]